



COURSE OUTLINE

The course description is online @ <http://camosun.ca/learn/calendar/current/web/chem.html>

⚡ Please note: the College electronically stores this outline for five (5) years only.
It is **strongly recommended** you keep a copy of this outline with your academic records.
You will need this outline for any future application/s for transfer credit/s to other colleges/universities.

1. Instructor Information

(a)	Instructor:	Steve McKinnon		
(b)	Office Hours:	After classes or by appointment		
(c)	Location:	F348A		
(d)	Phone:	370-3472	Alternative Phone:	
(e)	Email:	mckinnons@camosun.bc.ca		
(f)	Website:	D2L		

2. Intended Learning Outcomes

Upon completion of this course the student will be able to:

1. Utilize the specialized vocabulary and nomenclature based on the IUPAC system of organic compounds to name and draw structures for many simple organic compounds containing the common functional groups.
2. Write chemical reactions to illustrate numerous transformations between organic functional groups.
3. Draw structural and stereoisomers of organic compounds and name stereoisomers based upon the IUPAC system of nomenclature.
4. Demonstrate an understanding of the factors that influence the rate of a chemical reaction, deduce the rate of a chemical reaction from time/concentration data, and utilize rate laws to perform kinetic calculations.
5. Apply the laws of thermodynamics and account for the factors that lead to spontaneous physical and chemical changes.
6. Explain how and why reactions attain equilibrium positions and perform calculations pertaining to equilibrium systems.
7. Describe redox reactions, use electrochemical data to predict the spontaneity of redox reactions, and comprehend the structures of electrochemical cells.
8. Describe various acid-base theories and apply these theories to acid-base reactions in aqueous solution.
9. Perform experiments in the areas of preparative organic, preparative inorganic, physical and analytical chemistry and use the various associated pieces of laboratory equipment.

3. Required Materials

- (a) **Mastering Chemistry Course Code.** \$50 Access Code can be purchased from the Pearson Website <http://www.pearson.com.au/9781442563902> or the bookstore. If you choose to purchase a new textbook or ebook (see below) then a Mastering Chemistry Course Code is included. (valid for 24 months after activation).
- (b) Lab Experiments: Chemistry 120 Laboratory Manual, (In-house)
- (c) Safety Glasses

Other Recommended Materials

CHEMISTRY, The Central Science: a Broad Perspective[®] by Brown, Lemay, Bursten, Langford, Sagatys, and Duffy. Prentice Hall, Australian 3rd edition.

\$145 (Hard copy with access code)

\$114 (ebook with access code)

The 1st and 2nd editions or other texts are also acceptable, but you will still need to purchase a **Mastering Chemistry access code.**

4. Course Content and Schedule

Subject	Material Covered	Classes (approximate)	Textbook chapters*
Organic Chemistry	Alkane /Alkenes structure and properties, including naming alkanes / alkenes, reactions and stereochemistry, functional groups and some reactions.	6	22 to 27. Selected topics.
Chemical Kinetics	Reaction rates, change in concentration with time, temperature and rate, reaction mechanisms and catalysis	3	15
Thermochemistry	Energy, 1 st law of thermodynamics, enthalpy, calorimetry, Hess' Law, enthalpies of formation	2	14
Thermodynamics	Spontaneity, 2 nd law of thermodynamics, entropy, Gibbs Free Energy, free energy and temperature, free energy and equilibrium	2	14
Equilibrium	Equilibrium constants, heterogeneous equilibria, working with equilibrium constants	3	16
Acids and Bases	Acids and bases, pH scale, Ka and Kb, auto-ionization of water, acid strength of ions	3	17
Aqueous equilibria	Titration, common ion effect, buffers, solubility equilibrium	1	18
Electrochemistry	Redox reactions, balancing redox equations, half cells and the Nernst equation	1	4, 19

*textbook chapters are from Brown, LeMay, Bursten; 3rd Australian edition

5. Basis of Student Assessment (Weighting)

- (a) Midterm Test I (15%) - Learning Outcomes 1-3
- (b) Midterm Test II (15%) - Learning Outcomes 4-5
- (c) Final Exam (30%) – Comprehensive
- (d) MyLab Mastering (15 %)
- (e) Laboratory (25%)
 - 1. Students must complete a minimum of 70% of the laboratory work to pass the laboratory component of Chem 120. Students must pass the laboratory portion (>50%) of the course in order to obtain credit for Chem 121.
 - 2. Test score that is not as high as that of the final exam will be dropped automatically and its weight redistributed to the final exam. For anyone who misses both tests, your final exam will then be 60% of the course grade.
 - 3. Students must write each test as scheduled, no exceptions.

Important Dates

Week

III May 19 (Fri): **Test I in Lab from 2:30 to 4:20 pm**

IV May 22 (Mon): Victoria Day – no school

V Jun 1 (Thurs): Last Day to Withdraw or Change to Audit...

V Jun 2 (Fri): **Test II in Lab from 2:30 to 4:20 pm**

Exams Jun 19-21: Exam Period for Summer 2017

6. Grading System

Standard Grading System (GPA)

Percentage	Grade	Description	Grade Point Equivalency
90-100	A+		9
85-89	A		8
80-84	A-		7
77-79	B+		6
73-76	B		5
70-72	B-		4
65-69	C+		3
60-64	C		2
50-59	D	Minimum level of achievement for which credit is granted; a course with a "D" grade cannot be used as a prerequisite.	1
0-49	F	Minimum level has not been achieved.	0

Temporary Grades

Temporary grades are assigned for specific circumstances and will convert to a final grade according to the grading scheme being used in the course. See Grading Policy E-1.5 at camosun.ca for information on conversion to final grades, and for additional information on student record and transcript notations.

Temporary Grade	Description
I	<i>Incomplete:</i> A temporary grade assigned when the requirements of a course have not yet been completed due to hardship or extenuating circumstances, such as illness or death in the family.
IP	<i>In progress:</i> A temporary grade assigned for courses that, due to design may require a further enrollment in the same course. No more than two IP grades will be assigned for the same course. (For these courses a final grade will be assigned to either the 3 rd course attempt or at the point of course completion.)
CW	<i>Compulsory Withdrawal:</i> A temporary grade assigned by a Dean when an instructor, after documenting the prescriptive strategies applied and consulting with peers, deems that a student is unsafe to self or others and must be removed from the lab, practicum, worksite, or field placement.

7. Recommended Materials or Services to Assist Students to Succeed Throughout the Course

LEARNING SUPPORT AND SERVICES FOR STUDENTS

There are a variety of services available for students to assist them throughout their learning. This information is available in the College calendar, at Student Services, or the College web site at camosun.ca.

STUDENT CONDUCT POLICY

There is a Student Conduct Policy **which includes plagiarism**. It is the student's responsibility to become familiar with the content of this policy. The policy is available in each School Administration Office, at Student Services, and the College web site in the Policy Section.

Spring 2017 - Preliminary Lab Schedule:

Chem 121-001 - Wednesdays and Fridays, 9:30 - 12:20 pm in Fisher 354

<u>Week/ Starts on</u>	<u>Wednesday</u>	<u>Friday</u>
1 May 1 st	Lab Orientation - attendance mandatory	Exp. 1 - Preparation of Acetylsalicylic acid
2 May 8 th	Group A Exp. 3 - Preparation of Benzoic acid	Group B Exp. 3 - Preparation of Benzoic acid
3 May 15 th	Exp. 2 - Analysis of an unknown acid	Midterm I
4 May 21 st	Exp. 6 - The Rate of Iodination of Acetone	Lecture
5 May 29 th	Exp. 10 - Thermochemistry	Midterm II
6 June 5 th	Exp. 7 - Determination of pKa	Lecture
7 June 12 th	Exp. 9 - Preparation of CuCl	Exam Info & Review
Final exams - June 20 th - 22 nd		

Note: Dates are subject to change

All students are required to purchase their own **safety glasses**.

Safety glasses are mandatory for all lab activities.

The following are the general report guidelines for Chem 121-001. These are not required for all lab reports. Specific instructions will be given the day of each experiment.

Title of the experiment: State the title of the experiment you have just carried out.

Date: When you did the experiment.

Name: Your name and lab partner's name (if applicable)

Objective: State what you want to achieve by doing the experiment in one or two sentences. Be very brief and to the point.

Procedures: You can write the following: Please refer to Chem. 120 lab manual, 2009 Edition. pp. xx-xx. Record any changes to the given procedures.

Data: Organize any data, whether numerical or descriptive, in a **neat table** (or tables if applicable). Report such things as **unknown numbers**, **concentrations** of solutions, **masses** of reactants and products. Any relevant data recorded on a rough data sheet should be copied here. Do not forget to write **chemical equations** here.

Discussion and Calculations: In this part of the report, you will make sense out of the data you have obtained. If you obtain a product, calculate the **percentage yield**. Provide a physical description of your product. Answer any questions posed in the procedure. Show **all** the calculations you do, but there is no need to be repetitive. For example, if you perform 3 or 4 titrations using the same two solutions, then you only need to show the calculation for one trial. In cases where you have not obtained the results you were hoping for, provide a very brief explanation.

Conclusion: In **one or two sentences**, state your results.

* Lab reports should be written in **ink, including all calculations** or typed. If you are not using a computer to graph data, use graph paper.

* Lab reports are normally due at the time of the next experiment. You will be informed in advance if there are any changes to the due date. **Late reports will not be accepted.**

* Make sure to **staple** the pages of your report together, including any **rough data sheets**. You lose 1 (one) mark if the pages of your report are not stapled together.

Notes on graphing data:

*Data should be plotted on graph paper (1 mm square), or using a graphing program (MS Excel).

When using graph paper, make sure your plot takes up the entire page with an appropriate range for each axis.

When graphing with MS Excel, use the x,y-scatter format.

Titles - Make sure to include a descriptive title (what is the graph for?). Avoid using "vs" or "as a function of" in titles plot titles. For example, "Calibration curve for the determination of..."

Labels - the x and y axes need labels including the appropriate units (either as **/units** or **(units)**).

*Data points should be clearly shown on the graph, with a best-fit line drawn through them. Don't force your best fit line through the origin and don't "connect-the-dots".

*When extracting data from a plot, either interpolate with lines or use the equation of the line and clearly show your work with a sample calculation (example below).

Example: a graph of Beer's Law (Absorbance $\propto$ concentration) is linear,

in the form: **$y = mx + b$**

y is your dependent variable (Absorbance)

x is your independent variable (concentration)

Therefore, Absorbance = (m x Concentration) + b

If you know the absorbance of an unknown sample you can then figure out the concentration of that sample. Substitute the absorbance in for y and solve for x.

$(\text{Absorbance} - b) / m = \text{Concentration}$