



## COURSE OUTLINE

The course description is online @ <http://camosun.ca/learn/calendar/current/web/chem.html>

Ω Please note: the College electronically stores this outline for five (5) years only.  
It is **strongly recommended** you keep a copy of this outline with your academic records.  
You will need this outline for any future application/s for transfer credit/s to other colleges/universities.

### 1. Instructor Information

(a)	Instructor:	David Stuss, M.Sc.		
(b)	Office Hours:	Mon - Thurs 16:00 – 17:00 or by appointment		
(c)	Location:	Fisher F348D		
(d)	Phone:	(250) 370-3436	Alternative Phone:	
(e)	Email:	stussd@camosun.bc.ca		

### 2. Intended Learning Outcomes

(No changes are to be made to these Intended Learning Outcomes as approved by the Education Council of Camosun College.)

Upon completion of this course, the student will be able to:

1. Use dimensional analysis, metric and SI units in performing chemical calculations.
2. Utilize the specialized vocabulary and nomenclature of chemistry and name chemical compounds, and identify and construct chemical formulas.
3. Summarize the characteristics of electrons, protons and neutrons, and identify their roles as components of atoms, ions and isotopes, including radioisotopes.
4. Describe atomic structure, the differences between elements, and the role of the periodic table in organizing elements within a coherent theoretical and empirical system.
5. Describe and account for the periodic table trends concerning atomic number, atomic radius, ionization energy and electronegativity.
6. Compare the formation and characteristics of ionic and molecular compounds.
7. Perform mathematical calculations involving chemical formulas, molecular weights, moles, Avogadro's number and Molarity.
8. Balance chemical equations, including use of the mole concept, and solve stoichiometry problems.
9. Account for the general characteristics of the gas, liquid, and solid states.
10. Conduct experiments in basic chemistry, utilizing common chemistry laboratory equipment with an enhanced knowledge and practice in basic lab skills.

### 3. Required Materials

(a)	Text: Coursepack <b>(Mandatory)</b>	Chemistry 100 Course Notes, Lab Manual, and Problem Sets, 2015 Edition. Camosun College Publications.
(b)	Safety Glasses <b>(Mandatory)</b>	Bookstore has "UVEX" safety eyewear – please check if using others
(c)	Lab coat <b>(Recommended)</b>	Bookstore has cloth coats available – please check if using another type
(d)	Covered Footwear <b>(Mandatory)</b>	Exposed feet (e.g. sandals, flip-flops) are not permitted in the lab during experiments
(e)	Scientific Calculator <b>(Recommended)</b>	Available in bookstore. Smartphones / PDAs or similar devices cannot be used during testing or labs.

#### 4. Course Content and Schedule

##### Locations & Times

	Time	Location
<b>Lecture</b>	Monday 6:00 – 6:50 PM Wednesday 6:00 – 8:50 PM	Fisher Building, Room F334 Fisher Building, Room F334
<b>Lab</b>	Monday 7:00 – 8:50 PM	Fisher Building, Room F300

##### Lecture Plan

Unit	Topic	Unit	Topic
1	Measurements & Calculations	7	Chemical Bonding
2	Introductory Terminology	10	Organic Chemistry
3	Chemical Formulas & Names	8	Gases
4	Calculations Based Upon Formulas	9	Liquids & Solutions
5	Stoichiometry	11	Radioactivity
6	Periodic Table & Electron Distributions		

Lectures & homework exercises will follow the coursepack at a pace of approximately one unit per week.

##### Lab & Exam Schedule

Week	Lab Date	Lab No.	Lab Name
1	5 Sep	-	<i>Labour Day Holiday</i>
2	12 Sep	-	Safety Orientation (Mandatory)
3	19 Sep	1	Density
4	26 Sep	2	Identifying Liquids
5	3 Oct	3	Separating Mixtures
6	10 Oct	-	<i>Thanksgiving Holiday</i>
7	17 Oct	4	Heat of Combustion
<b>8</b>	<b>24 Oct</b>	-	<b>Midterm Exam</b>
9	31 Oct	5	Recycling Copper
10	7 Nov	6	The Fe / CuSO <sub>4</sub> Reaction
11	14 Nov	-	Molecular Model Tutorial
12	21 Nov	7	The Cu / AgNO <sub>3</sub> Reaction
13	28 Nov	11	The Magnesium HCl Reaction
14	5 Dec	-	Review / Problem Sets (Mandatory)

**Quizzes:** 6 in-class, 30 minute review quizzes on recent material

Quiz	Date**	Quiz	Date**
1	Sep 21	4	Nov 2
2	Oct 5	5	Nov 16
3	Oct 19	6	Nov 30

(\*Quiz dates may change with advance notice provided by the instructor. Quiz content will be discussed in class and posted online in D2L).

## 5. Basis of Student Assessment (Weighting)

Labs (8 x 2.5%)	20%
Quizzes (5/6)	20%
Midterm Exam	20%
Final Exam (comprehensive)	40%

1. To write the final exam you must achieve a minimum final score of **50%** on laboratory work.
2. You must pass **both** the lecture portion and the laboratory portion in order to pass the course.
3. In the theory section of the course, if the final exam mark is higher than the combined midterm + quiz mark, it will replace the combined midterm + quiz mark.
4. Quizzes are worth 4% each. The worst quiz score will be dropped when calculating the final quiz mark.
5. There will be no make-up quizzes or midterm exam. The weight of a missed quiz / midterm will be reassigned to the final exam.

### The Laboratory Mark

The lab mark is based on participation and the laboratory report. A student that participates in a laboratory class without completing the lab report will receive a minimum score of 50% on that lab (a report must still be submitted).

Students must watch an introductory Lab Safety video in the first lab class before they can begin any experiments. In the event of missing the Lab Safety presentation students are responsible for watching the (30 minute) safety DVD, available from the technician's office prior to their first experiment.

Wearing of **safety goggles** is **mandatory** in all labs. Students who forget safety goggles will not be allowed to complete the lab.

**Punctual attendance in all the lab periods is mandatory.** There are **NO EXCEPTIONS** other than an official doctor's note. Missed labs without adequate reasons will result in a mark of zero for that lab. **Permissions for an exception must be documented by email permission from the instructor and by a submitting the doctor's note.**

Laboratory reports can usually be completed in-class but are otherwise due one week late (i.e. the following lab period). The lab manual has been designed to allow students to hand in the completed pages taken directly from the manual. Each lab partner must hand in a separate report even if though lab partners are expected to share equally in experimental work.

## 6. Grading System

*(No changes are to be made to this section unless the Approved Course Description has been forwarded through the Education Council of Camosun College for approval.)*

### Standard Grading System (GPA)

Percentage	Grade	Description	Grade Point Equivalency
90-100	A+		9
85-89	A		8
80-84	A-		7
77-79	B+		6
73-76	B		5
70-72	B-		4
65-69	C+		3
60-64	C		2
50-59	D	Minimum level of achievement for which credit is granted; a course with a "D" grade cannot be used as a prerequisite.	1
0-49	F	Minimum level has not been achieved.	0

## Temporary Grades

Temporary grades are assigned for specific circumstances and will convert to a final grade according to the grading scheme being used in the course. See Grading Policy E-1.5 at [camosun.ca](http://camosun.ca) for information on conversion to final grades, and for additional information on student record and transcript notations.

Temporary Grade	Description
I	<i>Incomplete:</i> A temporary grade assigned when the requirements of a course have not yet been completed due to hardship or extenuating circumstances, such as illness or death in the family.
IP	<i>In progress:</i> A temporary grade assigned for courses that, due to design may require a further enrollment in the same course. No more than two IP grades will be assigned for the same course. (For these courses a final grade will be assigned to either the 3 <sup>rd</sup> course attempt or at the point of course completion.)
CW	<i>Compulsory Withdrawal:</i> A temporary grade assigned by a Dean when an instructor, after documenting the prescriptive strategies applied and consulting with peers, deems that a student is unsafe to self or others and must be removed from the lab, practicum, worksite, or field placement.

## 7. Recommended Materials or Services to Assist Students to Succeed Throughout the Course

### LEARNING SUPPORT AND SERVICES FOR STUDENTS

There are a variety of services available for students to assist them throughout their learning. This information is available in the College calendar, at Student Services, or the College web site at [camosun.ca](http://camosun.ca).

### STUDENT CONDUCT POLICY

There is a Student Conduct Policy **which includes plagiarism**. It is the student's responsibility to become familiar with the content of this policy. The policy is available in each School Administration Office, at Student Services, and the College web site in the Policy Section.