



## COURSE OUTLINE

The course description is online @ <http://camosun.ca/learn/calendar/current/web/chem.html>

Ω Please note: the College electronically stores this outline for five (5) years only.  
It is **strongly recommended** you keep a copy of this outline with your academic records.  
You will need this outline for any future application/s for transfer credit/s to other colleges/universities.

### 1. Instructor Information

(a)	Instructor:	David Stuss, M.Sc.		
(b)	Office Hours:	Mon-Thurs 14:30 – 15:20 <i>or by appointment</i>		
(c)	Location:	Fisher F348D		
(d)	Phone:	(250) 370-3436	Alternative Phone:	
(e)	Email:	stussd@camosun.bc.ca		

### 2. Intended Learning Outcomes

(No changes are to be made to these Intended Learning Outcomes as approved by the Education Council of Camosun College.)

Upon completion of this course the student will be able to:

1. Identify, describe and account for the general characteristics of gases, liquids and solids - interionic and intermolecular forces; vaporization and condensation; melting and freezing; specific characteristics of water.
2. Utilize solution terminology, account for and compare the solubilities of ionic and molecular compounds, and describe the impact of temperature and pressure on solubility.
3. Describe the characteristics of solubility equilibria and use mathematical techniques employed in dealing with this phenomenon.
4. Describe and account for the colligative and osmotic properties of aqueous solutions.
5. Account for differences in the rates of chemical reactions, apply Le Chatelier's Principle to equilibrium processes, and explain how catalysts influence reaction rates.
6. Apply mathematics and equilibrium constant expressions to descriptions of reversible reactions and chemical equilibria.
7. Identify Arrhenius, Bronsted and Lewis acids and bases, and describe the chemical properties of each type of substance.
8. Describe the ionization of water, the pH scale, weak and strong acids and bases, neutralization and the actions of buffer solutions.
9. Perform mathematical calculations involving pH, hydronium ion concentrations and acid-base titrations.
10. Define oxidation and reduction and assign oxidation numbers to the elements of substances involved in oxidation-reduction reactions. Demonstrate the ability to use oxidation numbers in balancing redox reactions.
11. Demonstrate an understanding of electrochemistry and account for the characteristics and uses of the standard hydrogen electrode, standard reduction potentials, electrolytic and voltaic cells.
12. Describe the characteristics of the major types of organic compounds – alkanes, alkenes, alkynes, aromatic hydrocarbons, alcohols, ethers, aldehydes and ketones, carboxylic acids and esters, amines and amides.

### 3. Required Materials

(a)	Textbook <b>(Strongly Recommended)</b>	"Chemistry, The Central Science" by Brown, LeMay & Bursten, 3 <sup>rd</sup> Australian Edition.
(b)	Lab Manual <b>(Mandatory)</b>	CHEM 110 Lab Manual, Camosun College Publications.
(c)	Safety Glasses <b>(Mandatory)</b>	Bookstore has "UVEX" safety eyewear – please check if using others
(d)	Lab coat <b>(Optional)</b>	Bookstore has cloth coats available – please check if using another type
(e)	Covered Footwear <b>(Mandatory in Lab)</b>	Exposed feet (e.g. sandals, flip-flops) are not permitted in the lab
(f)	Scientific Calculator <b>(Mandatory)</b>	Smartphones / PDAs or similar devices cannot be used during testing or labs.

All required materials are available at the Camosun College Bookstore.

### 4. Course Content and Schedule

<b>Credits</b>	4 credits	<b>Number of weeks</b>	14
<b>Workload / week</b>	3 h lecture 3 h lab 6 h study	<b>Pre-requisite</b>	CHEM 100 or Chemistry 11

#### Locations & Times

	Time	Location
<b>Lecture</b>	Thursday 6:30 – 9:20 PM	Fisher Building, Room F334
<b>Lab</b>	Tuesday 6:30 – 9:20 PM	Fisher Building, Room F354

#### Lecture Plan

Unit	Topic	Unit	Topic
1	States of Matter	6	Acid-Base Chemistry
2	Thermochemistry	7	Oxidation & Reduction
3	Solutions & Solubility	8	Electrochemistry
4	Chemical Kinetics	9	Organic Compounds
5	Chemical Equilibrium		

#### Important Dates

Date	Event
Tuesday Oct. 7 <sup>th</sup>	Term Test #1
Tuesday Nov. 4 <sup>th</sup>	Term Test #2
Tuesday Nov. 11 <sup>th</sup>	Remembrance Day
December 8 <sup>th</sup> – 16 <sup>th</sup>	Final Exam Week (Exam date TBA)

## Lab & Test Schedule

Week Number	Lab Date (Tuesdays)	Experiment No.	Experiment
1	Sep 2 <sup>nd</sup>	-	Review Lecture
2	Sep 9 <sup>th</sup>	-	Lab Safety / Orientation (Mandatory)
3	Sep 16 <sup>th</sup>	4	Precipitation reactions
4	Sep 23 <sup>th</sup>	6	Acid-Base Titrations
5	Sep 30 <sup>th</sup>	1	Energy Changes
6	<b>Oct 7<sup>th</sup></b>	-	<b>Term Test 1</b>
7	Oct 14 <sup>th</sup>	2	Reaction rates
8	Oct 21 <sup>nd</sup>	3	Equilibria
9	Oct 28 <sup>th</sup>	7	Vitamin C, Aspirin, Magnesia
10	<b>Nov 4<sup>th</sup></b>	-	<b>Term Test 2</b>
11	<b>Nov 11<sup>th</sup></b>	-	<b>Remembrance Day</b>
12	Nov 18 <sup>th</sup>	8	Acid-Base Titration Curves
13	Nov 25 <sup>th</sup>	10	Oxidation-Reduction Reactions
14	Dec 2 <sup>nd</sup>	-	Lecture / Review

## Quiz Schedule

Quizzes will be held at the beginning of **each lab period** (beginning in the second week of class). There will be no quizzes on midterm dates (Oct. 8<sup>th</sup> and Nov 4<sup>th</sup>).

## 5. Basis of Student Assessment (Weighting)

Labs (8 x 2.5%)	20%
Quizzes (10 x 1%)	10%
Midterm Tests (2 x 15%)	30%
Final Exam (comprehensive)	40%

- To write the final exam you must achieve a minimum final score of **50%** on laboratory work.
- You must pass **both** the lecture portion and the laboratory portion in order to pass the course.
- Weekly quizzes will be based on homework questions assigned in the previous week.
- There will be no make-up midterm tests or quizzes. The weight of a missed quiz / midterm will be reassigned to the final exam.
- If the percent score on the final exam mark is higher than the combined midterm mark, it will be used for the final mark on the theory section of the course (70% of the course mark).

## The Laboratory Mark

- Each lab has 2 components, the Pre-lab assignment and the Lab Report.
- Pre-Lab assignments can be found in the lab manual, and can be completed after reading through the lab protocol. They must be submitted at the **beginning** of the lab period. Late pre-labs will receive a late penalty (-10% per day).
- Lab Report worksheets will be provided online (in D2L) for students to print off prior to the lab period. Lab reports can usually be completed in-class but are otherwise due the following lab period (i.e., one week later). Lab partners must hand in their own separate reports even if though lab partners are expected to share equally in experimental work.
- Wearing of **safety goggles** is **mandatory** in all labs. Students who forget safety goggles will **not be allowed** to complete the lab. Students will have the option to store labeled goggles with the instructor in between labs. Covered footwear is also required.
- Attendance** in **all** the lab periods is **mandatory**. There are no exceptions other than an official doctor's note. Missed labs without adequate reasons will result in a mark of zero for that lab. *Permissions for an exception must be documented by email permission from the instructor and by a submitting the doctor's note.*

## 6. Grading System

(No changes are to be made to this section unless the Approved Course Description has been forwarded through the Education Council of Camosun College for approval.)

### Standard Grading System (GPA)

Percentage	Grade	Description	Grade Point Equivalency
90-100	A+		9
85-89	A		8
80-84	A-		7
77-79	B+		6
73-76	B		5
70-72	B-		4
65-69	C+		3
60-64	C		2
50-59	D	Minimum level of achievement for which credit is granted; a course with a "D" grade cannot be used as a prerequisite.	1
0-49	F	Minimum level has not been achieved.	0

### Temporary Grades

Temporary grades are assigned for specific circumstances and will convert to a final grade according to the grading scheme being used in the course. See Grading Policy E-1.5 at [camosun.ca](http://camosun.ca) for information on conversion to final grades, and for additional information on student record and transcript notations.

Temporary Grade	Description
I	<i>Incomplete:</i> A temporary grade assigned when the requirements of a course have not yet been completed due to hardship or extenuating circumstances, such as illness or death in the family.
IP	<i>In progress:</i> A temporary grade assigned for courses that, due to design may require a further enrollment in the same course. No more than two IP grades will be assigned for the same course. (For these courses a final grade will be assigned to either the 3 <sup>rd</sup> course attempt or at the point of course completion.)
CW	<i>Compulsory Withdrawal:</i> A temporary grade assigned by a Dean when an instructor, after documenting the prescriptive strategies applied and consulting with peers, deems that a student is unsafe to self or others and must be removed from the lab, practicum, worksite, or field placement.

## 7. Recommended Materials or Services to Assist Students to Succeed Throughout the Course

### LEARNING SUPPORT AND SERVICES FOR STUDENTS

There are a variety of services available for students to assist them throughout their learning. This information is available in the College calendar, at Student Services, or the College web site at [camosun.ca](http://camosun.ca).

### STUDENT CONDUCT POLICY

There is a Student Conduct Policy **which includes plagiarism**. It is the student's responsibility to become familiar with the content of this policy. The policy is available in each School Administration Office, at Student Services, and the College web site in the Policy Section.