



Chemistry 100
INTRODUCTORY CHEMISTRY

Section 003

Winter Semester 2012

COURSE OUTLINE

This course is a basic introduction to chemistry and is intended for students with little or no background in chemistry. Topics include chemical formulae and equations, simplest formula, atomic mass, mole concept, molarity, periodic table, molecules and chemical bonding, and some descriptive chemistry. Experiments will emphasize basic lab techniques.

1. Instructor Information

- (a) Instructor Jamie Doran, Ph.D.
- (b) Office hours Monday, 1:30 to 2:20 pm
 Tuesday, 11:30 to 1:20 pm
 Wednesday, 1:30 to 2:20 pm
 Thursday, 1:30 to 2:20 pm
 Friday, 1:30 to 2:20 pm

Everyone is welcome whenever my office door is open.

Appointments may be made to meet at times other than those listed above.

Office hours will be extended prior to exams.

- (c) Location Room 350C, Fisher Building, Lansdowne Campus
- (d) Phone 250.370.3441
- (e) E-mail jdoran@camosun.bc.ca

2. Intended Learning Outcomes

At the end of this course, students will be able to:

- Utilize the specialized vocabulary and nomenclature of chemistry.
- Use metric and SI units in performing chemical calculations.
- Describe the experimental discovery of subatomic particles, summarize the characteristics of electrons, protons and neutrons, and identify their roles as components of atoms.
- Communicate an understanding of atomic structure, the differences between elements, and the role of the periodic table in organizing elements within a coherent theoretical and empirical system.
- Describe and account for the periodic table trends concerning atomic number, atomic radius, ionization energy and electronegativity.
- Demonstrate an ability to name chemical compounds, and identify and construct chemical formulas.
- Compare the formation and characteristics of ionic and molecular compounds.
- Demonstrate an ability to perform mathematical calculations involving chemical formulas, molecular weights, moles, Avogadro's number and molarity.
- Balance chemical equations, demonstrate an understanding of the information they provide chemists and solve stoichiometry problems.
- Identify and account for the general characteristics of the gas state and solve mathematical problems involving Boyle's Law, Charles' Law, and Avogadro's Law.
- Communicate an understanding of radioisotopes, nuclear fission and nuclear fusion.

3. Required Materials

(a) Course Text & Laboratory Manual

Chemistry 100 Study Notes, Supplementary Problems, & Laboratory Manual, 2011 Edition. Camosun College Publications.

(b) General Materials and Supplies

Safety glasses Safety glasses are required when handling hazardous chemicals.

Students are required to provide their own pair of safety glasses.

Students lacking safety glasses will not be permitted in the laboratory when experiments are being conducted.

Lab coats Lab coats are required for conducting experiments in this course.

Students are required to provide their own lab coats. Students lacking lab coats will not be permitted to work in the laboratory when experiments are being conducted.

Scientific calculator Calculators may be required in the lab, in lectures and during exams. Each student is required to provide her or his own calculator.

4. Course Content and Schedule

	Credits	4 credits
holding session	In-class workload	4 h of lectures per week. 2 h lab period per week. Experiments are conducted in most lab periods as outlined in the schedule below. This lab period will also be used for the midterm exam, and for a review prior to the final exam.
	Out-of-class workload	6 hours per week
	Number of weeks	14 weeks
	Pre-requisite	Principles of Math 10, or Foundations of Math & Pre-calculus 10, or MATH 053, or MATH 057; or assessment.

Course Times and Locations

Lectures	Monday 8:30 to 9:20 AM Fisher Building, Room F360
	Wednesday 8:30 to 9:20 AM Fisher Building, Room F360
	Thursday 8:30 to 9:20 AM Fisher Building, Room F360
	Friday 8:30 to 9:20 AM Fisher Building, Room F360
Laboratory Periods	Tuesday 8:30 to 10:20 AM Fisher Building, Room F300

Please see the laboratory and midterm exam schedule below.

Lecture Outline

A detailed outline of the lecture material is provided in the Table of Contents of *Chemistry 100 Notes*. Notably, this book has been designed specifically for this course to present many relevant examples of the chemistry of life and the environment including those intended to stimulate interest and curiosity.

1. Measurements and Calculations: SI & other scientific units; SI prefixes; metric conversions; measurements, scientific notation, & significant figures; density calculations; calculations involving energy changes.

2. Introductory Terminology: scientific method; physical & chemical changes; elements, compounds and mixtures; metals and nonmetals; atoms and molecules; protons, neutrons and electrons; ions and isotopes; atomic masses.

3. Chemical Formulas and Names: composition of chemical compounds; formulas and naming of molecular compounds; meaning of ionic formulas and naming of ionic compounds; compounds containing polyatomic ions; formulas and names of acids.

4. Calculations Based Upon Formulas: molecular mass; formula mass; percentage composition; the mole; grams to moles and moles to grams conversions; moles of molecular of ionic compounds; Avogadro's Number.

5. Stoichiometry: balancing chemical equations; stoichiometry - problems based upon chemical equations; limiting reactant calculations; percentage yield calculations; calculations involving exothermic or endothermic chemical reactions.

6. Periodic Table and Electron Distributions: chemical families; electron levels and orbitals (sublevels); electron distribution in atoms; electron dot formulae; trends in atomic radii (size), ionization energies & chemical reactivity.

7. Chemical Bonding: formation of ionic compounds; formation of molecular compounds; electron dot formula representations; electronegativity and bond polarity; molecular geometry and polarity.

8. Gases: general nature of gases; factors affecting gas volume; Boyle's Law - gas pressure & volume; absolute temperature scale; Charles' Law - gas temperature & volume; STP standard conditions of gas temperature and pressure ; molar gas volume; partial pressures of gases; gases and diving; gas stoichiometry.

9. Liquids and Solutions: general properties of liquids; hydrogen bonding; vapour pressure and boiling point; solubility; solution concentration & diluting solutions; electrolytes, dissociation equations & ion concentrations in solution; pH scale; solution stoichiometry.

10. Organic Chemistry: why so many organic compounds?; structural formulas and isomers; naming of hydrocarbons & alcohols; optional: addition and substitution reactions in organic chemistry.

11. Radioactivity: Radioactive substances; alpha, beta & gamma rays & associated decay; optional: production of radioisotopes; half-life and dating; medical applications.

Laboratory & Midterm Exam Schedule

Please familiarize yourself in advance with the lab practices and safety information presented on pages 5 & 6 of the Laboratory Manual.

Week 1. Tuesday, January 10 th	Chemistry Laboratory & Safety Orientation Session
Week 2. Tuesday, January 17 th	Experiment 1. Density.
Week 3. Tuesday, January 24 th	Experiment 3. Separating Mixtures.
Week 4. Tuesday, January 31 st	Experiment 4. Heat of Combustion.
Week 5. Tuesday, February 7 th	Experiment 15. Accuracy & Precision Experiment 5 (start). Recycling Copper.
Week 6. Tuesday, February 14 th	Experiment 5. Recycling Copper.
Week 7. Tuesday, February 21 st	Experiment 5 (completion). Recycling Copper. Experiment 7. The Copper - Silver Nitrate Reaction.
Week 8. Tuesday, February 28 th	Midterm Exam
Week 9. Tuesday, March 6 th	Experiment 8. Water of Hydration.
Week 10. Tuesday, March 13 th	Molecular modeling: geometry & polarity.
Week 11. Tuesday, March 20 th	Experiment 10. Volume of a Gas.
Week 12. Tuesday, March 27 th	Experiment 12. Neutralization.
Week 13. Tuesday, April 3 rd	Experiment 13. Synthesis of Aspirin
Week 14. Tuesday, April 10 th	Final Exam Review Period

5. Basis of Student Assessment (Weighting)

(a) Laboratory Reports

Attendance in the lab periods is mandatory. No laboratory experiment can be missed without an acceptable reason submitted in writing such as a suitable note from a Medical Doctor. Laboratory reports are due in the following experimental lab period unless otherwise stated. The lab manual has been designed to allow students to hand in the completed pages taken directly from the manual. A formal laboratory report may be required for one designated experiment. Each lab partner must hand in a separate report even if though lab partners typically share equally in experimental work. The value the lab reports contribute to the final grade is **15%**.

(b) Quizzes

These will compose **25%** of the final grade. There will be five quizzes each of equal value (5%).

- Quiz 1. Chapters 1 & 2
- Quiz 2. Chapters 3 & 4
- Quiz 3. Chapter 5
- Quiz 4 Chapters 6 & 7
- Quiz 5 Chapters 8 & 9

Quizzes will be typically scheduled a few days to a week following the completion of lectures concerning the Chapter(s) from which material is to be tested. Attempts will be made to schedule quizzes on days when students do not have other tests or exams.

There is no quiz on Chapter 10 or 11 material. The last material to be presented in the semester would be tested on the final exam.

(c) Midterm Exam

This exam will compose **25%** of the final grade, and will test material from Chapters 1 to 5. It is scheduled for Wednesday, October 28th in the period normally used for a lab experiment. This exam will need to be written in either F300 - the lab room. Please note that the date for this test will be confirmed in class as the pace of the course with this particular class becomes evident, and efforts are made to avoid having this exam scheduled on the same day as students' other exams.

If any quiz or test is missed due to illness or any other justifiable reason, a student may either take a substitute test scheduled at a mutually agreeable time, or choose to add the percentage value of that test that of the final exam.

(d) Final Exam

The final exam is a comprehensive exam that will cover all of the material presented in the lecture portion of the course with an emphasis on material that follows Chapter 3. The value this exam contributes to the final grade is **35 %**.

The time and location of the Chem 060 final exam will be published by the College during the Fall Semester.

Attendance at the final exam is mandatory. Appropriate documentation must accompany an explanation for absence.

6. Grading System

Standard Grading System (GPA)

Percentage	Grade	Description	Grade Point Equivalency
90-100	A+		9
85-89	A		8
80-84	A-		7
77-79	B+		6
73-76	B		5
70-72	B-		4
65-69	C+		3
60-64	C		2
50-59	D		1
0-49	F	Minimum level has not been achieved.	0

Temporary Grades

Temporary grades are assigned for specific circumstances and will convert to a final grade according to the grading scheme being used in the course. See Grading Policy at camosun.ca or information on conversion to final grades, and for additional information on student record and transcript notations.

Temporary Grade	Description
I	<i>Incomplete:</i> A temporary grade assigned when the requirements of a course have not yet been completed due to hardship or extenuating circumstances, such as illness or death in the family.
IP	<i>In progress:</i> A temporary grade assigned for courses that are designed to have an anticipated enrollment that extends beyond one term. No more than two IP grades will be assigned for the same course.
CW	<i>Compulsory Withdrawal:</i> A temporary grade assigned by a Dean when an instructor, after documenting the prescriptive strategies applied and consulting with peers, deems that a student is unsafe to self or others and must be removed from the lab, practicum, worksite, or field placement.

Temporary grades are assigned for specific circumstances and will convert to a final grade according to the grading scheme being used in the course. See Grading Policy E-1.5 at camosun.ca for information on conversion to final grades, and for additional information on student record and transcript notations.

7. Recommended Materials or Services to Assist Students to Succeed Throughout the Course

Students will find that the *Chemistry 100 Study Notes* is very well designed to mirror the lecture content of this course. Importantly, this book provides very many practice questions and their answers, as does the *Chemistry 100 Supplemental Questions* that is included in the course package. These practice questions are representative of those that will appear on the quizzes, the midterm exam and the final exam.

LEARNING SUPPORT AND SERVICES FOR STUDENTS

There are a variety of services available for students to assist them throughout their learning. This information is available in the College Calendar, Registrar's Office or the College web site at <http://www.camosun.bc.ca>

Please Note:

Students may not use recording devices in the classroom without the prior permission of the instructor. However, the instructor's permission is not required when the use of a recording device is sanctioned by the College's Resource Centre for Students with Disabilities in order to accommodate a student's disability and when the instructor has been provided with an instructor notification letter which specifies the use of a recording device. Recordings made in the classroom are for the student's personal use only, and distribution of recorded material is prohibited.

ACADEMIC CONDUCT POLICY

There is an Academic Conduct Policy. It is the student's responsibility to become familiar with the content of this policy. The policy is available in each School Administration Office, Registration, and on the College web site in the Policy Section.

www.camosun.bc.ca/divisions/pres/policy/2-education/2-5.html