



School of Arts & Science
CHEMISTRY AND GEOSCIENCE DEPARTMENT

CHEM 224-01
Introduction to Analytical Chemistry
2011F

COURSE OUTLINE

This course provides an introduction to the theory and practice of chemical separation and identification. The major techniques are examined including chromatography, solvent extraction, spectroscopy and electrochemistry. The problems associated with data acquisition and the statistics of sampling are stressed throughout.

Ω Please note: this outline will be electronically stored for five (5) years only. It is strongly recommended students keep this outline for your records.

Prerequisite: Chem 121 (C grade minimum)

1. Instructor Information

(a)	Instructor:	Blair Surridge		
(b)	Office Hours:	Thurs: 10:30 – 12:20 Wed & Fri: 1:30-3:20		
(c)	Location:	F350A		
(d)	Phone:	370-3438	Alternative Phone:	
(e)	Email:	SurridgeB@camosun.bc.ca bsurridge@shaw.ca (home)		

2. Intended Learning Outcomes

(No changes are to be made to this section, unless the Approved Course Description has been forwarded through EDCO for approval.)

At the end of this course, the student will possess an enhanced ability to:

1. Define and calculate the mean, median, mode, variance and standard deviation for a series of replicate analyses. Estimate the population mean from analysis of a small number of trials. Test for the rejection or retention of suspect data. Explain and use the least squares procedure to graph experimental data.
2. Describe and explain the procedures for gravimetric and titrimetric analyses: obtain data that falls within the established margins of error for the methods.
3. Derive and apply the Beer-Lambert law and use internal and external standards to ensure the validity of the analysis. Distinguish between absorption, emission, fluorescence and phosphorescence. Obtain absorption and emission spectra from various sources and perform a complete quantitative analysis on the samples provided. Explain and use light scattering techniques to estimate the turbidity of solutions.

- Distinguish between the major modes of radioactive decay and between the activity of the sample and the dose received by the absorber. Estimate the age of fossils and artifacts via carbon and argon dating techniques and the concentrations of trace materials using neutron activation and isotope dilution techniques.
- Identify and describe the mode of operation for the four major types of electrode. Distinguish between constant current and constant potential coulometry and use them to estimate the concentrations of particular ions in solution. Distinguish between normal and pulsed polarography and analyze polarograms obtained from mixtures of metal ions.
- Describe, explain and apply the techniques of solvent extraction, distillation, sublimation, and the major forms of chromatography to the separation of a mixture.
- Discuss the basis for improvements in the signal to noise ratio of a measurement. Distinguish between the Fourier transform and continuous wave methods of recording data. Explain the process of analogue to digital conversion.
- Construct a null point hypothesis; use one or two tailed significance tests to reject or retain the hypothesis. Use a paired t test to compare two different methods of analysis for the same sample

3. Required Materials

Text	◆ "Quantitative Chemical Analysis" 8th Edition, by Daniel C. Harris (Freeman and Company)
Other	◆ Chem 224 Lab Manual (Safety glasses mandatory & lab coat recommended) ◆ A Small hard backed laboratory notebook (from bookstore)
In Library On Reserve	◆ "Fundamentals of Analytical Chemistry" 8 th addition, by Skoog, West, Holler, and Crouch

4. Course Content and Schedule (news items & supplemental learning material can be found on D2L)

Lectures:

Monday	10:30 to 11:20 am in Wilna Thomas Building, WT201
Wednesday	10:30 to 11:20 am in Wilna Thomas Building, WT201
Friday	10:30 to 11:20 am in Ewing Building, E201

Unit	Topic	Textbook Reference* (Select topics only)
1	Analytical process, measurement, experimental error, and statistics	Ch. 0, 1, 3, and 4
2	Classical methods (Gravimetric and Titration)	Ch. 26

3	Methods of Calibration and Quality Assurance	Ch.4 and 5
4	Spectrochemical Methods	Ch. 17, 18, 19, and 20
5	Methods of Separation	Ch. 21, 22, 23, and 24
6	Electrochemical Methods	Ch. 13 and 14

*- see class lectures for details

Chem. 224 Lab Schedule Monday 2:30-5:20pm in F356

(note: your lab group will be assigned to either schedule A or B as shown below to maximize time on lab instrumentation)

Week	Lab Date	Experiment No.	Schedule A	Schedule B
I	Sept 5 th	Labour Day, No meeting	-	-
II	Sept 12 th	Exp # 1, Introduction & skills assessment	Exp#1	Exp#1
III	Sept 19 th	Exp # 2, Analysis of halide ions using silver nitrate	Exp#2	Exp#2
IV	Sept 26 th	Exp # 3, Calibration of Instruments	Exp#3	Exp#11
V	Oct 3 th	Exp # 11, Isotopic dilution and separation of mixtures	Exp#11	Exp#3
VI	Oct 10 th	No Lab - Thanksgiving Day!!	-	-
VII	Oct 17 th	Midterm (2.5hrs)	-	-
VIII	Oct 24 th	Exp # 4, UV/Vis spectroscopy	Exp#4	Exp#9 (part1)
IX	Oct 31 st	Exp # 5, Atomic absorption spectroscopy	Exp#5	Exp#9 (part2)
X	Nov 7 th	Exp # 9, Chromatography (Part1)	Exp#9 (part1)	Exp#4
XI	Nov 14 th	Exp #9 Chromatography Part 2 (Analysis of BPA in Water)	Exp#9 (part2)	Exp#5
XII	Nov 21 st	Exp # 7, Ion selective electrodes	Exp#7	Exp#7
XIII	Nov 28 th	Tour of Federal Lab: Institute of Ocean Sciences	everyone	everyone
XIV	Dec 5 th	Lecture Wrap-up and Final Exam Review	-	-

5. Basis of Student Assessment (Weighting)

Labs	25%
Quizzes	15% (in class)*
Midterm (Units I, 2, 3, & 4**)	15% (Week VII Lab Period, 2-hour)
In class group presentation	10% (Week XIII)
Final Exam (comprehensive)	35% (TBA ~Week XV, 3 hours in Dec)

*Tentatively 6 quizzes scheduled. You will receive at least 3 days of notice before a quiz and details will be posted on D2L!!

**May cover on part of unit 4 material

Notes:

- (1) Student must pass the lab and lecture component of the course to obtain credit for Chem 224. All labs are to be attended and individual lab reports completed following the format provided in the detailed lab orientation handout.
- (2) Immediate contact must be made with instructor for missed labs due to illness or family emergencies for arrangements to be made. For more information see lab orientation handout.
- (3) A test score that is not as high as that of the December final exam will be dropped automatically and its weight redistributed to the final exam. For example, if the midterm is missed your final exam will then be 50% of the course grade!

- (4) No one is allowed to write exams late and there will be no exceptions. Early exam is a privilege and not a right; thus, at full discretion of the instructor. In the event that the midterm is missed the marks will be weighted to the final exam.

6. Grading System

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Standard Grading System (GPA)

Percentage	Grade	Description	Grade Point Equivalency
90-100	A+		9
85-89	A		8
80-84	A-		7
77-79	B+		6
73-76	B		5
70-72	B-		4
65-69	C+		3
60-64	C		2
50-59	D	Minimum level of achievement for which credit is granted; a course with a "D" grade cannot be used as a prerequisite.	1
0-49	F	Minimum level has not been achieved.	0

Temporary Grades

Temporary grades are assigned for specific circumstances and will convert to a final grade according to the grading scheme being used in the course. See Grading Policy E-1.5 at camosun.ca for information on conversion to final grades, and for additional information on student record and transcript notations.

Temporary Grade	Description
I	<i>Incomplete:</i> A temporary grade assigned when the requirements of a course have not yet been completed due to hardship or extenuating circumstances, such as illness or death in the family.
IP	<i>In progress:</i> A temporary grade assigned for courses that, due to design may require a further enrollment in the same course. No more than two IP grades will be assigned for the same course. (For these courses a final grade will be assigned to either the 3 rd course attempt or at the point of course completion.)
CW	<i>Compulsory Withdrawal:</i> A temporary grade assigned by a Dean when an instructor, after documenting the prescriptive strategies applied and consulting with peers, deems that a student is unsafe to self or others and must be removed from the lab, practicum, worksite, or field placement.

7. Important Dates

Week

III	Sept. 20: Fee deadline
VI	Oct.10: Thanksgiving Monday-College Closed
X	Nov. 11 (Friday): Remembrance Day—College Closed
XIV	Exam Period for Winter 2011

Use this link to check out scholarships and bursaries

<http://camosun.ca/learn/calendar/current/pdf/financial-assistance.pdf>

8. Recommended Materials or Services to Assist Students to Succeed Throughout the Course

LEARNING SUPPORT AND SERVICES FOR STUDENTS

There are a variety of services available for students to assist them throughout their learning. This information is available in the College calendar, at Student Services or the College web site at camosun.ca.

STUDENT CONDUCT POLICY

There is a Student Conduct Policy **which includes plagiarism**. It is the student's responsibility to become familiar with the content of this policy. The policy is available in each School Administration Office, at Student Services and on the College web site in the Policy Section.

Please Note:

Students may not use recording devices in the classroom without the prior permission of the instructor. However, the instructor's permission is not required when the use of a recording device is sanctioned by the College's Resource Centre for Students with Disabilities in order to accommodate a student's disability and when the instructor has been provided with an instructor notification letter which specifies the use of a recording device. Recordings made in the classroom are for the student's personal use only, and distribution of recorded material is prohibited. Cell phones should be turned off will in class.